

Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals _____
- b. metal ions _____
- c. anodes _____
- d. cathodes _____
- e. half-cell reactions: Zn _____
- f. spontaneous overall reactions _____

- g. half-cell reduction potentials _____
- h. E_{cell} _____

i. Write overall cell equation: $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$ (10 marks)
Write half-reactions with phases: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$ and $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$

Voltaic Cells Study Guide

JG Myers



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