

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $n = 1000$, $m = 1000$

(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$
 (C) $1 \text{ mol H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

2. An acid-base reaction involves hydrogen ions. Suppose instead you consider hydroxide ions. If 1.0 g of acid were neutralized by base, how much base would you need? (1.0 g of acid, 1.0 g of base)

(A) $n = 1000$, $m = 1000$

(B) $1 \text{ g H}_2\text{O}$, $1 \text{ g H}_2\text{O}$

3. An acid-base reaction involves H^+ ions. H^+ ions are the smallest cations and anions. What is the ratio of H^+ ions to OH^- ions?

(A) $1:1$, (B) $1:2$, (C) $2:1$, (D) $2:2$

4. Which is the correct combination of properties for this following equation?



(A) 1 , (B) 2 , (C) 3 , (D) 4

5. A solution of acetic acid and its conjugate base is prepared with a pH of 4.0. What is the ratio of acetic acid to its conjugate base?



(A) 10^4 , (B) 10^{-4} , (C) 10^2 , (D) 10^{-2} , (E) 10^0

6. An acid-base equilibrium involves acetic acid (CH_3COOH) and its conjugate base (CH_3COO^-). After this reaction was completed, resulting with a solution with a pH of 4.0. What is the ratio of acetic acid to its conjugate base?



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Anthony S. Fauci



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