

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. What is the pH of the solution? (correct)
- (A) 2.0      (B) 3.0      (C) 4.0      (D) 5.0      (E) 6.0

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

3. The pH of a 0.1 M acetic acid solution is 3.0. What is the concentration of acetate ions? (correct)
- (A) 0.01 M      (B) 0.02 M      (C) 0.05 M      (D) 0.1 M      (E) 0.2 M

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

4. Which of the following is a weak acid? (correct)
- (A) HCl      (B) HNO<sub>3</sub>      (C) H<sub>2</sub>SO<sub>4</sub>      (D) H<sub>3</sub>PO<sub>4</sub>      (E) H<sub>2</sub>O

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

5. A 0.1 M solution of a weak acid has a pH of 3.0. What is the concentration of the conjugate base? (correct)
- (A) 0.01 M      (B) 0.02 M      (C) 0.05 M      (D) 0.1 M      (E) 0.2 M

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (G)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (H)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

6. The pH of a 0.1 M acetic acid solution is 3.0. What is the concentration of acetate ions? (correct)
- (A) 0.01 M      (B) 0.02 M      (C) 0.05 M      (D) 0.1 M      (E) 0.2 M

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

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**Chao Zhang**



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