

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, amount of substance
 (D) volume of hydrogen, amount of substance, amount of substance, amount of substance

(A) $n = 1000$, $m = 1000$

(B) $n = 10^{-3}$, $m = 10^{-3}$, $p = 10^{-3}$, $q = 10^{-3}$ (C) $n = 10^{-3}$, $m = 10^{-3}$, $p = 10^{-3}$, $q = 10^{-3}$
 (D) $n = 10^{-3}$, $m = 10^{-3}$, $p = 10^{-3}$, $q = 10^{-3}$ (E) none of these are correct

2. An acid-base indicator has a K_a of 1.0×10^{-5} . In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $n = 10^{-5}$, $m = 10^{-5}$, $p = 10^{-5}$, $q = 10^{-5}$

(B) $n = 10^{-5}$, $m = 10^{-5}$, $p = 10^{-5}$, $q = 10^{-5}$ (C) $n = 10^{-5}$, $m = 10^{-5}$, $p = 10^{-5}$, $q = 10^{-5}$

3. An acid-base indicator has a K_a of 1.0×10^{-5} . In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $n = 10^{-5}$, $m = 10^{-5}$, $p = 10^{-5}$, $q = 10^{-5}$ (B) $n = 10^{-5}$, $m = 10^{-5}$, $p = 10^{-5}$, $q = 10^{-5}$

4. Which of the following is a weak acid? (correct)



(E) H_2O (F) H_2O (G) H_2O (H) H_2O

5. A solution of an acid has a pH of 3.0. The concentration of the acid is 0.1 mol/L. What is the K_a of the acid? (correct)



(E) H_2O (F) H_2O (G) H_2O (H) H_2O

6. An acid-base indicator has a K_a of 1.0×10^{-5} . In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)



Multiple Choice Stoichiometry Test

Roman Wölfel



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