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 $2 h_2 + 1 o_2 \rightarrow 2 h_2o$
 $pb + oh \rightarrow 2 hcl + 2 h_2o + 1 pbcl_2$
 $albr_3 + 3 k_2so_4 \rightarrow 6 kbr + 1 al_2 so_4 + 3 ch_4 + 2 o_2$
 $1 co_2 + 2 h_2o + c_3h_8 + 5 o_2 \rightarrow 3 co_2 + 4 h_2o + c_8h_{18}$
 $25 o_2 + 16 co_2 + 18 h_2o$

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 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas
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 $2 nacl + f_2 \rightarrow 2 naf + cl_2$
 $4 h_2 + 1 o_2 \rightarrow 2 h_2o$
 $5 pb + oh \rightarrow 2 hcl + 2 h_2o + 1 pbcl_2$
 $6 albr_3 + 3 k_2so_4 + 6 kbr + 1 al_2so_4 + 3 ch_4 + 2 o_2$
 $1 co_2 + 2 h_2o + 8 c_3h_8 + 5 o_2 \rightarrow 3 co_2 + 4$

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