

Chem 216: Pre-Lab Example

The Pre-Lab assignment is simply a means to show that you are prepared for the experiment for the day. Should you struggle in the preparation of the Pre-Lab or with any of the calculations for each lab feel free to email your GSI or Dr. Koreeda. Your GSI will also go over any details at the beginning of the lab period.

Your Pre-Lab should be written in you lab notebook before you come into the lab. Some variation off this example is fine. Come up with a system that works for you, but the basic details should still be there. If you would like to include an outline of the procedure, that would be appropriate as well. The Pre-Lab is also a good launching point for writing your lab report.

Acetylation of Aniline & Identification of Unknown Aniline Derivative

The conversion of aniline or other amines to their corresponding acetamide allows for the protection of the amine functionality from attack by electrophiles as well as oxidizing agents. As a result of adding the protecting group, a crystalline solid is generally afforded that may be easily purified through recrystallization. Once purified the solid then can be characterized using simple lab techniques like IR spectroscopy, melting-point analysis, or the Beilstein test to verify proper identity and synthesis.



Equivalents	Compound	MW, d, M	Moles	Amount
1.0 eq	aniline	93.1 g/mol	5.4 mmol	500 mg
1.0 eq	conc. HCl	37% w/w, 36.5 g/mol	5.5 mmol	0.45 mL
1.2 eq	acetic anhydride	102 g/mol 1.08 g/mL	6.4 mmol	0.60 mL
1.2 eq	sodium acetate	82.0 g/mol	6.5 mmol	530 mg

When using concentrated HCl, the solution from the bottle is generally 37% HCl by weight of solution. So, with a density of 1.2 g/mL,

$$0.45 \text{ mL conc. HCl} \times \frac{1.2 \text{ g HCl}}{1 \text{ mL}} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1000 \text{ mmol}}{1 \text{ mol}} \times 0.37 = 5.5 \text{ mmol HCl}$$

Experimental Organic Chemistry Pre Lab

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Velavancorresponding**



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