

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are the characteristic properties of a weak acid? _____
- (A) pKa and molar conductivity (B) pKa and pH
- (C) molar conductivity (D) molar conductivity and pH

2. Which of the following is CORRECT? In a weak acid solution, the concentration of _____ is _____.
- (A) H^+ is $10^{-4} mol\ dm^{-3}$ (B) H^+ is $10^{-2} mol\ dm^{-3}$ (C) H^+ is $10^{-6} mol\ dm^{-3}$ (D) H^+ is $10^{-8} mol\ dm^{-3}$
- (E) H^+ is $10^{-4} mol\ dm^{-3}$ (F) H^+ is $10^{-2} mol\ dm^{-3}$ (G) H^+ is $10^{-6} mol\ dm^{-3}$ (H) H^+ is $10^{-8} mol\ dm^{-3}$

3. An indicator is a weak acid. In a solution of a weak acid, the concentration of the indicator is $1.0 \times 10^{-4} mol\ dm^{-3}$. The pH of the solution is 4.0. What is the pKa of the indicator? _____
- (A) 4.0 (B) 3.0 (C) 5.0 (D) 2.0 (E) 6.0

4. An indicator is a weak acid. In a solution of a weak acid, the concentration of the indicator is $1.0 \times 10^{-4} mol\ dm^{-3}$. The pH of the solution is 4.0. What is the pKa of the indicator? _____
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7. An indicator is a weak acid. In a solution of a weak acid, the concentration of the indicator is $1.0 \times 10^{-4} mol\ dm^{-3}$. The pH of the solution is 4.0. What is the pKa of the indicator? _____
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8. An indicator is a weak acid. In a solution of a weak acid, the concentration of the indicator is $1.0 \times 10^{-4} mol\ dm^{-3}$. The pH of the solution is 4.0. What is the pKa of the indicator? _____
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