

COURSE OF MAINTENANCE
 ELECTRICAL ENGINEERING DEPARTMENT

**LESSON 1:
 FUNDAMENTALS OF ELECTRICITY**

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 Subject Instructor

Atomic Structure

- An atom is the smallest particle of an element that retains the characteristics of that element.
 - Electron - negatively charged particle
 - Proton - positively charged particle
 - Neutron - uncharged particle
- Each type of atom has a certain number of electrons and protons that distinguishes it from the atoms of all other elements.
 - Atomic Number** - equals the number of protons
 - Atomic Weight** - equals the weight of the nucleus

Atomic Structure

- Shell** - the orbit in which an electron revolves
- Valence Electron** - an electron that is present in the outermost shell of an atom
- Ion** - an atom or a group of atoms with a net charge
- The number of electrons in each shell follows a predictable pattern according to the formula, $2N^2$, where N is the shell number

Categories of Materials

- Conductors**
 - materials that readily allow current
 - have a large number of free electrons and are characterized by one to three valence electrons in their structures
- Semiconductors**
 - placed between the conductors in their ability to carry current
 - have four valence electrons in their atomic structures
- Insulators**
 - poor conductors of electric
 - very few free electrons
 - characterized by more than four valence electrons in their atomic structures

Electrical Charge

- Electrical charge** - an electrical property of matter that is the result of an excess or deficiency of electrons, it is symbolized by Q
- Static electricity** is the presence of the net positive or negative charge on a material
- Electric field** - a force that acts between charges

Electrical Charge

- Electrical charge (Q) is measured in coulombs, symbolized by C
- One coulomb is the total charge possessed by 6.25×10^{18} electrons
- A single electron has a charge of $1.6 \times 10^{-19} C$

$$Q = \text{number of electrons} \times 1.6 \times 10^{-19} \text{ electrons/C}$$

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Ian Pickup



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