

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $n = 1000$ ,  $m = 1000$

(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$   
 (C)  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

2. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A)  $10^{10}$   
 (B)  $10^9$   
 (C)  $10^8$   
 (D)  $10^7$

(E)  $10^6$   
 (F)  $10^5$   
 (G)  $10^4$   
 (H)  $10^3$   
 (I)  $10^2$   
 (J)  $10^1$

3. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A)  $10^1$   
 (B)  $10^2$   
 (C)  $10^3$   
 (D)  $10^4$

4. Which of the following is a dimensional quantity? (correct)

(A)  $n = 1000$ ,  $m = 1000$

(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

5. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A)  $10^{10}$   
 (B)  $10^9$   
 (C)  $10^8$   
 (D)  $10^7$

(E)  $10^6$   
 (F)  $10^5$   
 (G)  $10^4$   
 (H)  $10^3$   
 (I)  $10^2$   
 (J)  $10^1$

6. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of pH 5.5, the indicator is present in the form of its conjugate base. What is the ratio of the concentration of the conjugate base to the concentration of the indicator? (correct)

(A)  $10^{10}$   
 (B)  $10^9$   
 (C)  $10^8$   
 (D)  $10^7$

# Multiple Choice Stoichiometry Test

**Paul Foglino**



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