

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH from the concentration of either ion.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution  |
| 4. hydroxide ion  | 5. hydronium ion            | 6. autoionization |
|                   | 7. acid solution            | 8. pH             |

## Key Equations

- |  |   |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (page 712)        | 2. $\text{pOH} = -\log[\text{OH}^-]$ (page 712) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (page 712) | 4. $\text{pH} + \text{pOH} = 14.00$ (page 712)  |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Write the number in the space provided.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (page 712)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (page 712)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (page 712)
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9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (page 712)
10.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (page 712)

1. conjugate acid
2. conjugate base
3. conjugate acid
4. conjugate base
5. conjugate acid
6. conjugate base
7. conjugate acid
8. conjugate base
9. conjugate acid
10. conjugate base

# 19 2 Hydrogen Ions And Acidity Answer Key

**AW Chickering**



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