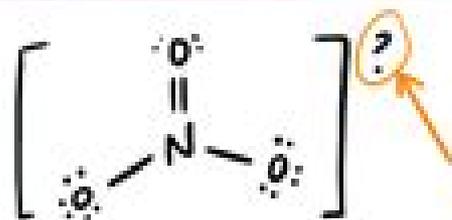


# [ FORMAL CHARGE ]<sup>+/-</sup>

$$F.C. = (\text{valence } e^-) - (\text{nonbonding } e^-) - \frac{1}{2} (\text{bonding } e^-)$$

- \* The amount of electrons in a specific atom when the molecule shares electrons equally.
- \* Only applies to covalent molecules.

## Examples



Overall charge comes from addition of all atomic formal charges!

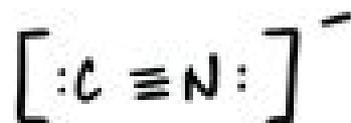
Q: What is the formal charge of N?

A:

$$F.C. = 5 - 0 - (\frac{1}{2} \times 8)$$

$$F.C. = 5 - 4$$

$$F.C. = +1$$



Q: What is the formal charge of carbon?

$$A: F.C. = 4 - 2 - (\frac{1}{2} \times 6)$$

$$F.C. = 4 - 2 - 3$$

$$F.C. = -1$$

# Formal Charges

**C Cleary**



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## **Table of Contents Formal Charges**

1. Understanding the eBook Formal Charges
  - The Rise of Digital Reading Formal Charges
  - Advantages of eBooks Over Traditional Books
2. Identifying Formal Charges
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Formal Charges
  - User-Friendly Interface
4. Exploring eBook Recommendations from Formal Charges
  - Personalized Recommendations
  - Formal Charges User Reviews and Ratings
  - Formal Charges and Bestseller Lists
5. Accessing Formal Charges Free and Paid eBooks
  - Formal Charges Public Domain eBooks

- Formal Charges eBook Subscription Services
- Formal Charges Budget-Friendly Options
- 6. Navigating Formal Charges eBook Formats
  - ePub, PDF, MOBI, and More
  - Formal Charges Compatibility with Devices
  - Formal Charges Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Formal Charges
  - Highlighting and Note-Taking Formal Charges
  - Interactive Elements Formal Charges
- 8. Staying Engaged with Formal Charges
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Formal Charges
- 9. Balancing eBooks and Physical Books Formal Charges
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Formal Charges
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Formal Charges
  - Setting Reading Goals Formal Charges
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Formal Charges
  - Fact-Checking eBook Content of Formal Charges
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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