

# Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a) Anode: \_\_\_\_\_
- b) Cathode: \_\_\_\_\_
- c) Anion: \_\_\_\_\_
- d) Cation: \_\_\_\_\_
- e) Half-reaction at anode: \_\_\_\_\_
- f) Half-reaction at cathode: \_\_\_\_\_
- g) Net redox reaction: \_\_\_\_\_
- h) Salt: \_\_\_\_\_
- i) Write overall cell equation with phases.  $\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu(s)}$  (10 marks)

\_\_\_\_\_

# Voltaic Cells Study Guide

**Christian G. Meyer**



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