

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R + \dots$$

The equipartition value $C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R$ is reached when the temperature is high enough so that $T \gg \theta_M$ for all the modes. However, if the temperature is lower, some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower. The heat capacity is lower because some of the modes are not active and the heat capacity is lower.

Solutions to exercises

The heat capacity of a diatomic gas is given by $C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R + \nu_V^\ddagger R$. When equipartition is valid, $\nu_R^\ddagger = 2$ and $\nu_V^\ddagger = 1$, so $C_{V,m} = \frac{7}{2}R$.

$$C_{V,m} = \frac{7}{2}R = \frac{7}{2} \times 8.314 \text{ J K}^{-1} \text{ mol}^{-1} = 29.12 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 298 K, $T < \theta_R$ and $T < \theta_V$, so only translational and rotational modes are active.

$$C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R = \frac{3}{2}R + 2R = \frac{7}{2}R = 29.12 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 1000 K, $T > \theta_R$ and $T < \theta_V$.

$$C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R + \nu_V^\ddagger R = \frac{3}{2}R + 3R + R = \frac{8}{2}R = 33.26 \text{ J K}^{-1} \text{ mol}^{-1}$$

At 2000 K, $T > \theta_R$ and $T > \theta_V$.

$$C_{V,m} = \frac{3}{2}R + \nu_R^\ddagger R + \nu_V^\ddagger R = \frac{7}{2}R + R = \frac{9}{2}R = 37.40 \text{ J K}^{-1} \text{ mol}^{-1}$$

Physical Chemistry 9th Solution Manual

JR Anderson



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