

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$   
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1V_1=P_2V_2$   
Isolate for P<sub>2</sub>:  $P_2=(P_1V_1)/V_2$   
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$   
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1=34.75 + 273=307.75 \text{ K}$   
 $T_2=78.41 + 273=351.41 \text{ K}$   
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
Get rid of P  
 $V_1/T_1=V_2/T_2$   
Isolate V<sub>2</sub>:  $V_2=(V_1/T_1)T_2$   
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV=nRT$  will work. (n= moles and R= 0.08206 L atm/K mol)  
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$  OOPS we need to convert from grams to moles!,  $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$   
n: Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$   
Isolate V:  $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Final Review Guide Chemistry

**Michal Rosen-Zvi**



## **2nd Semester Final Review Guide Chemistry:**

## Unveiling the Energy of Verbal Beauty: An Mental Sojourn through **2nd Semester Final Review Guide Chemistry**

In a global inundated with screens and the cacophony of instant communication, the profound energy and emotional resonance of verbal art frequently fade into obscurity, eclipsed by the constant barrage of sound and distractions. However, nestled within the musical pages of **2nd Semester Final Review Guide Chemistry**, a captivating perform of literary splendor that pulses with natural thoughts, lies an memorable trip waiting to be embarked upon. Composed by a virtuoso wordsmith, that magical opus instructions visitors on a mental odyssey, delicately exposing the latent possible and profound impact stuck within the elaborate internet of language. Within the heart-wrenching expanse with this evocative analysis, we shall embark upon an introspective exploration of the book is key subjects, dissect their captivating writing type, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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