

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. K_a | 6. K_b |
| | 7. weak base | |

Key Equations

- | | |
|---|---------------------------------------|
| 1. $K_w = [H_3O^+][OH^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[H_3O^+][A^-]}{[HA]}$ |
| 3. $pH = -\log [H_3O^+]$ | 4. $pOH = -\log [OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Complete the following reactions.

1. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$
 Identify the conjugate acid-base pair of the reactants and products of this reaction.
 Conjugate acid: H_2SO_4 Conjugate base: HSO_4^-
2. $H_2SO_4 + H_2O \rightleftharpoons SO_4^{2-} + H_3O^+$
 Identify the conjugate acid-base pair of the reactants and products of this reaction.
 Conjugate acid: H_2SO_4 Conjugate base: SO_4^{2-}
3. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$
 Identify the conjugate acid-base pair of the reactants and products of this reaction.
 Conjugate acid: H_2SO_4 Conjugate base: HSO_4^-
4. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$
 Identify the conjugate acid-base pair of the reactants and products of this reaction.
 Conjugate acid: H_2SO_4 Conjugate base: HSO_4^-

5. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$
6. $H_2SO_4 + H_2O \rightleftharpoons SO_4^{2-} + H_3O^+$
7. $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$
8. $CH_3COOH + H_2O \rightleftharpoons CH_3COOH_2^+ + OH^-$
9. $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$
10. $CH_3COOH + H_2O \rightleftharpoons CH_3COOH_2^+ + OH^-$
11. $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$
12. $CH_3COOH + H_2O \rightleftharpoons CH_3COOH_2^+ + OH^-$
13. $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$
14. $CH_3COOH + H_2O \rightleftharpoons CH_3COOH_2^+ + OH^-$

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