

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $n = 1000$, $m = 1000$

(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$
 (C) $1 \text{ mol H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

2. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $\text{pH} = 4.5$, $\text{pH} = 4.5$, $\text{pH} = 4.5$, $\text{pH} = 4.5$

(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

3. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

4. Which of the following is a strong acid? (correct)



(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

5. A solution of an acid has a pH of 2.0. The concentration of the acid is 0.1 mol/L. What is the pK_a of the acid? (correct)



(B) $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $1 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$, $10 \text{ g H}_2\text{O}$, $1 \text{ mol H}_2\text{O}$

6. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)



Multiple Choice Stoichiometry Test

Leopold E. Klopfer



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