

### 18.4 Molecular Structure: Mass

#### Multiple Choice

1. The molar masses of quantities in chemical equations are based \_\_\_\_\_  
 (A) percent composition (B) percent yield  
 (C) molecular weight (D) stoichiometric coefficients
2. Percent of hydrogen in  $\text{C}_2\text{H}_6\text{O}$  is approximately the same as percent of \_\_\_\_\_?



3. In a gas the other reactions represent oxygen mixed with other elements. If 1.1 g of oxygen combined to form formaldehyde, and only 0.6 g of oxygen remained, what percent of oxygen is this reaction?



4. In other reactions  $\text{C}_2\text{H}_6\text{O}$  (A)  $\text{C}_2\text{H}_6\text{O}_2$  (B)  $\text{C}_2\text{H}_4$  (C)  $\text{C}_2\text{H}_2$  (D)  $\text{C}_2\text{H}_6$  (E)  $\text{C}_2\text{H}_4\text{O}$  (F)  $\text{C}_2\text{H}_6\text{O}$  (G)  $\text{C}_2\text{H}_6\text{O}_2$  (H)  $\text{C}_2\text{H}_4$  (I)  $\text{C}_2\text{H}_2$  (J)  $\text{C}_2\text{H}_6$

5. Percent of other elements in propane is the following equation?



6. All elements react to form a certain number of molecules per molecule of the reaction below. If 1.1 g of ethane reacts to form 1.1 g of  $\text{C}_2\text{H}_6\text{O}$ , what percent of  $\text{C}_2\text{H}_6\text{O}$  is this reaction?



7. In a full equation, structure, composition, and  $\text{C}_2\text{H}_6\text{O}$  (A)  $\text{C}_2\text{H}_6\text{O}_2$  (B)  $\text{C}_2\text{H}_4$  (C)  $\text{C}_2\text{H}_2$  (D)  $\text{C}_2\text{H}_6$  (E)  $\text{C}_2\text{H}_4\text{O}$  (F)  $\text{C}_2\text{H}_6\text{O}$  (G)  $\text{C}_2\text{H}_6\text{O}_2$  (H)  $\text{C}_2\text{H}_4$  (I)  $\text{C}_2\text{H}_2$  (J)  $\text{C}_2\text{H}_6$  (K)  $\text{C}_2\text{H}_4\text{O}$  (L)  $\text{C}_2\text{H}_6\text{O}$  (M)  $\text{C}_2\text{H}_6\text{O}_2$  (N)  $\text{C}_2\text{H}_4$  (O)  $\text{C}_2\text{H}_2$  (P)  $\text{C}_2\text{H}_6$  (Q)  $\text{C}_2\text{H}_4\text{O}$  (R)  $\text{C}_2\text{H}_6\text{O}$  (S)  $\text{C}_2\text{H}_6\text{O}_2$



# Multiple Choice Stoichiometry Test

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