

The Uncertainty Principle

- If the wavefunction is Ae^{ikx} , then the particle it describes has a definite state of linear momentum $px = +kh$.
- The position of the particle described by this wavefunction is completely unpredictable.
- In other words, if the momentum is specified precisely, it is impossible to predict the location of the particle.
- This statement is one-half of a special case of the Heisenberg uncertainty principle, one of the most celebrated results of quantum mechanics:

If the momentum is specified precisely, it is impossible to predict the location of the particle

Existence Principle

Andrew Bernstein



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