

Procedure #	Liquid or Solid Used	Reagent	Observation	Ppt contains	Centrifugate Contains	Applicable Equations
1	Group I, II, III, IV	HCl	White & Yellow	AgCl, Hg ₂ Cl ₂ , PbCl ₂	Group 2, 3, and 4	$Ag^+ + HCl \rightarrow AgCl + H^+$ $2Hg^{2+} + 2HCl \rightarrow Hg_2Cl_2 + 2H^+$ $Pb^{2+} + 2HCl \rightarrow PbCl_2 + 2H^+$
2	PPT From #1	Hot H ₂ O	PPT Still Present	AgCl, Hg ₂ Cl ₂	Pb ²⁺	$PbCl_2 + H_2O + Heat \rightarrow Pb^{2+}(aq) + Cl^-(aq)$
3	Half of decant from #2	K ₂ CrO ₄	Yellow Precipitate	PbCrO ₄	N/A	$Pb^{2+}(aq) + K_2CrO_4 \rightarrow PbCrO_4 + 2K^+$
	Other half of decant	H ₂ SO ₄	White Precipitate	PbSO ₄	N/A	$H_2SO_4 + Pb^{2+} \rightarrow PbSO_4 + 2H^+$
4a	PPT From #2	NH ₃	Black/Gray PPT	Hg ⁺ HgNH ₂ Cl	Ag(NH ₂) ⁺	$AgCl + 2NH_3 \rightarrow Ag(NH_2)^+ + Cl^-$ $Hg_2Cl_2 + NH_3 \rightarrow Hg(l) + HgNH_2Cl(s) + NH_4^+ + Cl^-$
4b	Decantate	HNO ₃	White Precipitate	AgCl	N/A	$Ag(NH_2)^+(aq) + 2H^+(aq) \rightarrow Ag^+(aq) + 2NH_4^+(aq)$ $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$
5a	Decant from #1	H ₂ O ₂ , (NH ₄) ₂ S, NH ₄ C ₂ H ₃ O ₂ , H ₂ O, Heat	Dark/Hunter Green	PbS, Bi ₂ S ₃ , CuS, CdS	Group III and IV	$Pb^{2+} + S^{2-} \rightarrow PbS$ $2Bi^{3+} + 3S^{2-} \rightarrow Bi_2S_3$ $Cu^{2+} + S^{2-} \rightarrow CuS$ $Cd^{2+} + S^{2-} \rightarrow CdS$
5b						
6a	PPT From #5	HNO ₃ , H ₂ SO ₄ , Heat	Yellow Crystal Forms, ions dissolve, and White Smoke Appears	Sulfur	Group II ions	$PbS + 2NO_3^- + 4H^+ \rightarrow Pb^{2+} + S + 2NO_2 + 2H_2O$ $CuS + 2NO_3^- + 4H^+ \rightarrow Cu^{2+} + S + 2NO_2 + 2H_2O$ $CdS + 2NO_3^- + 4H^+ \rightarrow Cd^{2+} + S + 2NO_2 + 2H_2O$ $Bi_2S_3 + 6NO_3^- + 12H^+ \rightarrow 2Bi^{3+} + 3S + 6NO_2 + 6H_2O$
6b	Decant from 6a	Ice, NH ₃ , HC ₂ H ₃ O ₂ , KCrO ₄	White Precipitate, Then Dissolve, Then Yellow PPT	PbSO ₄	Bi ³⁺ , Cu ²⁺ , Cd ²⁺	$Pb^{2+} + SO_4^{2-} + ICE \rightarrow PbSO_4 + Heat$ $PbSO_4 + 3NH_3 + C_2H_3O_2 \rightarrow Pb(C_2H_3O_2)_2 + 3NH_4^+ + SO_4^{2-}$ $Pb(C_2H_3O_2)_2 + KCrO_4 \rightarrow PbCrO_4 + 3[C_2H_3O_2]^- + K^+$
7a	Decant from #6	NH ₃				$Bi^3 + 3NH_3 + 3H_2O \rightarrow Bi(OH)_3 + 3NH_4^+$ $Cu^{2+} + 4NH_3 \rightarrow Cu(NH_3)_4^{2+}$ (Blue) $Cd^{2+} + 4NH_3 \rightarrow Cd(NH_3)_4^{2+}$
7b		PPT From #7a	H ₂ O, NaOH, SnCl ₂	Jet Black	Bi(S)	N/A

General Chemistry With Qualitative Analysis

Karin Nielsen-Saines



General Chemistry With Qualitative Analysis:

This book delves into General Chemistry With Qualitative Analysis. General Chemistry With Qualitative Analysis is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into General Chemistry With Qualitative Analysis, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to General Chemistry With Qualitative Analysis
- Chapter 2: Essential Elements of General Chemistry With Qualitative Analysis
- Chapter 3: General Chemistry With Qualitative Analysis in Everyday Life
- Chapter 4: General Chemistry With Qualitative Analysis in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of General Chemistry With Qualitative Analysis. This chapter will explore what General Chemistry With Qualitative Analysis is, why General Chemistry With Qualitative Analysis is vital, and how to effectively learn about General Chemistry With Qualitative Analysis.

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