

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

2. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. The pH of the solution is (correct)
- (A) 1.0      (B) 2.0      (C) 3.0      (D) 4.0      (E) 5.0

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

3. The pH of a 0.1 M acetic acid solution is 3.0. The concentration of acetate ions is (correct)
- (A) 0.01 M      (B) 0.02 M      (C) 0.03 M      (D) 0.04 M      (E) 0.05 M

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

4. Which of the following is a weak acid? (correct)
- (A)  $\text{HCl}$       (B)  $\text{H}_2\text{SO}_4$       (C)  $\text{H}_2\text{CO}_3$       (D)  $\text{HNO}_3$

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

5. A 0.1 M acetic acid solution has a pH of 3.0. The concentration of acetate ions is (correct)
- (A) 0.01 M      (B) 0.02 M      (C) 0.03 M      (D) 0.04 M      (E) 0.05 M

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

(E)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (G)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (H)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

6. In a 0.1 M acetic acid solution, the concentration of acetate ions is 0.01 M. The pH of the solution is (correct)
- (A) 1.0      (B) 2.0      (C) 3.0      (D) 4.0      (E) 5.0

(A)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (C)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{[\text{H}_2\text{O}]}$

# Multiple Choice Stoichiometry Test

**Anthony S. Fauci**



**Multiple Choice Stoichiometry Test:**

## Decoding **Multiple Choice Stoichiometry Test**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Multiple Choice Stoichiometry Test**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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