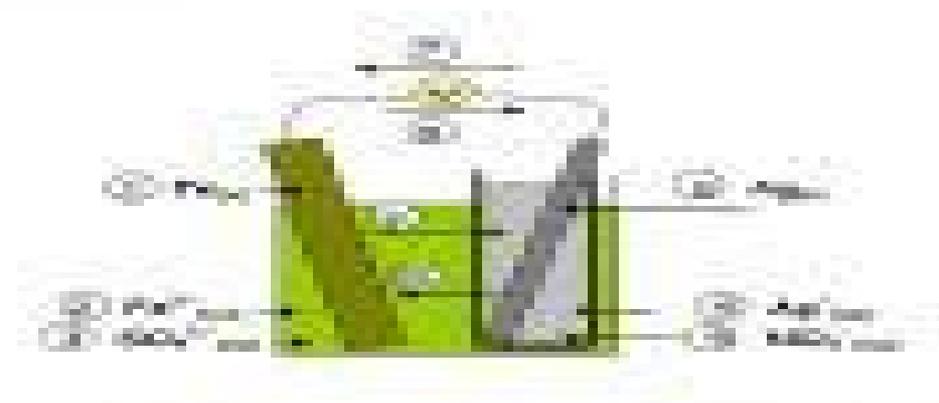


# Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. Anode: \_\_\_\_\_
  - b. Cathode: \_\_\_\_\_
  - c. Anion: \_\_\_\_\_
  - d. Cation: \_\_\_\_\_
  - e. Half-cell reactions: Zn: \_\_\_\_\_
  - f. Spontaneous overall reaction: \_\_\_\_\_
  - g. Write overall balanced cell: \_\_\_\_\_
  - h. Salt: \_\_\_\_\_
  - i. Write overall cell process:  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$  (10 marks)
- Write a spontaneous redox reaction for copper & silver.

\_\_\_\_\_

# Voltaic Cells Study Guide

**Ying Liu**



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