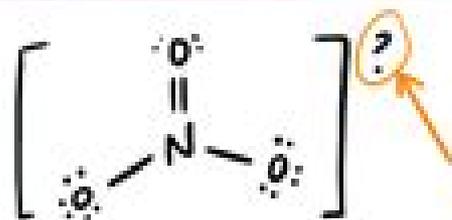


[FORMAL CHARGE]^{+/-}

$$F.C. = (\text{valence } e^-) - (\text{nonbonding } e^-) - \frac{1}{2} (\text{bonding } e^-)$$

- * The amount of electrons in a specific atom when the molecule shares electrons equally.
- * Only applies to covalent molecules.

Examples



Overall charge comes from addition of all atomic formal charges!

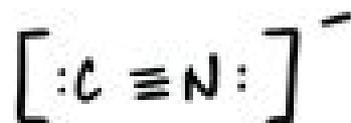
Q: What is the formal charge of N?

A:

$$F.C. = 5 - 0 - (\frac{1}{2} \times 8)$$

$$F.C. = 5 - 4$$

$$F.C. = +1$$



Q: What is the formal charge of carbon?

$$A: F.C. = 4 - 2 - (\frac{1}{2} \times 6)$$

$$F.C. = 4 - 2 - 3$$

$$F.C. = -1$$

Formal Charges

M Woodhall



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