

18.4 Molecular Structure: Mass

Multiple Choice:

1. The molar masses of quantities in chemical equations are based _____
 (A) percent composition (B) percent yield
 (C) molecular weight (D) stoichiometric coefficients
2. Percent of hydrogen in $\text{C}_2\text{H}_6\text{O}$ is _____ of the total mass of the compound?

(A) 15.0% (B) 12.5%

(C) 15.0% (D) 12.5% (E) 20.0% (F) percent of the atoms

3. In a gas the other reactions represent oxygen mixed with other elements. If 1.0 g of gas was produced by the reaction, and only produced 0.5 g of gas, what is the percent yield of sulfur in this reaction?

(A) 50.0% (B) 100.0% (C) 200.0% (D) 50.0%

(E) 100.0% (F) 200.0% (G) 50.0% (H) 200.0%

4. In this reaction $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$, what is the percent yield of oxygen?
 (A) 100% (B) 50% (C) 200% (D) 50%

5. What is the percent yield of oxygen in the following equation?



(A) 100% (B) 50% (C) 200% (D) 50%

6. Aluminum reacts with oxygen to produce aluminum oxide. If 1.0 g of aluminum is used and 0.5 g of aluminum oxide is produced, what is the percent yield?



(A) 100% (B) 50% (C) 200% (D) 50% (E) 200%

7. In a full experiment, structure-molecular weight of H_2SO_4 is what percent yield of sulfur, H_2SO_4 . After the reaction was completed, the yield was percent of the sulfur in the reaction mixture. What is the percent yield?



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RM Cervero

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